**Key Vocabulary: Week 3.0: Chemical Bonding and Formulas**

***Complete ALL Week 3.0 and 3.5 flashcards and study over the weekend.***

 ***Week 3.0- Flashcards on Monday AND Tuesday, Quiz on Tuesday, 7/2.***

1. **Lewis Dot Structures:** Diagram showing the valence electrons of an atom or ion.
2. **Covalent Bond:** Chemical bond in which electrons are shared. (Non-Metal/Non-Metal)
3. **Ionic Bond:** Chemical bond in which electrons are given and taken. (Metal/Non-Metal)
4. **Metallic Bond:** Chemical bond in which electrons flow freely. (Metal/Metal)
5. **Crystal:** The regular and repeating pattern of ions within an ionic solid.
6. **Electronegativity:** Idea that some atoms want to take electrons more than others.
7. **Representative Particles:** the simplest part of a substance that can represent all of that substance. Molecules of H2O are representative of water.
8. **Formula Unit:** Name given to 1 particle of an ionic compound. Ex. NaCl
9. **Molecule:** Name given to 1 particle of a covalent compound. Ex. CO2
10. **Polyatomic Ion:** Charged groups of atoms, usually covalent. Ex. SO42-
11. **NH41+ :** Ammonium
12. **CO32- :** Carbonate
13. **NO31- :** Nitrate
14. **SO42- :** Sulfate
15. **PO43- :** Phosphate
16. **OH1- :** Hydroxide