**Key Vocabulary: Week 4.5: Moles and Chemical Reactions**

***Complete ALL Week 4.0 and 4.5 flashcards and study over the weekend.***

 ***Week 4.5- Flashcards on Wednesday AND Thursday, Quiz on Thursday, 7/11.***

1. **Empirical Formula:** A molecular formula simplified to show the smallest whole number ratios of atoms. C6H12O6 → CH2O
2. **Gram Atomic Mass:** molar mass of an element or specific type of atom.
3. **Gram Formula Mass:** molar mass of a substance’s empirical formula.
4. **Gram Molecular Mass:** molar mass of an entire molecule.
5. **Representative Particles:** the simplest part of a substance that can represent all of that substance. Molecules of H2O are representative of water.
6. **Molecule:** a single particle of a covalent compound. Consists of 2 or more chemically-bonded atoms.
7. **Formula Unit:** the representative particle of an ionic crystal. Ex- NaCl.
8. **C3H8 :** propane (used in gas grills).
9. **H2CO3 :** carbonic acid (makes soda fizzy).
10. **H2O2 :** hydrogen peroxide (cleans cuts, disinfectant).
11. **Fe2O3 :** rust (reddish material found on corroded metals).
12. **C8H10N4O2 :** caffeine (stimulant found in coffee, energy drinks, etc).