**Key Vocabulary: Week 2.0: Gases and Solutions**

***Complete ALL Week 2.0 and 2.5 flashcards and study over the weekend.***

***Week 2.0- Flashcards on Monday AND Tuesday, Quiz on Tuesday, 6/25.***

1. **Pressure:** the force exerted by gas particles on each other and their containers. Usually exists as a "push" outwards.
2. **Volume:** the space taken up by a substance. Usually measured in liters (L).
3. **Temperature:** a measurement of the vibrations in matter due to energy. Higher temperatures = greater vibrations.
4. **Kelvin:** a temperature scale in which 0 is the point at which atoms no longer vibrate. This is known as "absolute zero" and is the lowest temperature possible. Used for "T" in the gas laws.
5. **Boyle's Law:** P1V1 = P2V2
6. **Charles' Law:** V1 / T1 = V2 / T2
7. **Ideal Gas Law:** PV = nRT (R is the gas constant, 0.0821)
8. **Combined Gas Law:** P1V1 / T1 = P2V2 / T2
9. **STP:** standard temperature and pressure. 273 K and 1 atm.
10. **Kinetic Molecular Theory:** describes gases as a large group of particles in constant random motion whose collisions produce pressure.
11. **Ideal Gas:** a simplified concept that describes gas particles as having no volume, and no attraction/repulsion with one another.
12. **Avogadro's Hypothesis:** states that the volume and moles of a gas are proportional. Allows us to use the molar volume assumption.
13. **Molar Volume:** assumption that 1 mol of an ideal gas occupies 22.4 L
14. **Dalton's Law of Partial Pressures:** the pressure of a mixture of gases is the sum of the pressures of each separate gas. Pt = P1 + P2 + P3 ......