**Key Vocabulary: Week 3.0: Reactivity**

***Complete ALL Week 3.0 and 3.5 flashcards and study over the weekend.***

***Week 3.0- Flashcards on Monday AND Tuesday, Quiz on Tuesday, 7/2.***

1. **Dissociation**: the idea that compounds (typically ionic) "break apart" to form separate ions in solution.
2. **Oxidation-Reduction Reaction:** a chemical reaction in which electrons are exchanged, changing the oxidation numbers of one or more reactants / products.
3. **Precipitation Reaction:** a chemical reaction in which dissolved ions (aq) react to produce a product that is fairly insoluble and "drops out" of solution as a solid (s).
4. **Law of Conservation of Mass:** “What goes in must come out.” The mass of all reactants entering a chemical reaction must match the mass of all products leaving it.
5. **Reactants:** Substances reacting or entering a reaction, left side.
6. **Products:** Substances made by a reaction, right side.
7. **Synthesis/Combination:** A + B 🡪 C Smaller substances combine to make a larger one.
8. **Decomposition:** C 🡪 A + B Larger substances break down to make smaller ones.
9. **Double Replacement:** AB + CD 🡪 AD + BC - A “double swap” where two reactants are rearranged by a chemical reaction.
10. **Single Replacement:** AB + C 🡪 AC +B - A “single swap” where a substance replaces another as part of a compound.
11. **Neutralization:** Acid/base reaction “X(OH) + HY 🡪 H2O + XY” where a substance carrying hydrogen (acid) reacts with a hydroxide base to make water and a salt, XY.
12. **Combustion:** A hydrocarbon reacts with oxygen to produce heat, carbon dioxide, and water. Burning!
13. **Oxidation Numbers:** The charge of an element within a compound.