**MON. Objective:** We will review the concept of ideal gases and the gas laws.

(6/24)**Product:**  I will solve for pressure, volume, or temperature using the gas laws.

**Exam or Quizzes**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ **Homework**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**TUE. Objective:** We will continue to review the concept of ideal gases and the gas laws.

(6/25) **Product:**  I will solve the ideal gas equation and utilize the concept of molar volume.

**Exam or Quizzes**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ **Homework**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**WED. Objective:** We will review solution chemistry and solubility rules.

(6/26)**Product:**  I will calculate the molarity of a solution.

**Exam or Quizzes**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ **Homework**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

 \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**THUR. Objective:** We will review and take our Exam over gases and solutions.

(6/27)**Product:**  I will pass Exam #2: Gases and Solutions.

**Exam or Quizzes**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ **Homework**: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

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**FRI. NO SCHOOL**

(6/28)

**Grades Received This Week:**

**\_\_\_\_\_ \_\_\_\_\_ \_\_\_\_\_ \_\_\_\_\_ \_\_\_\_\_ \_\_\_\_\_ \_\_\_\_\_**

**Parent Signature: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**Student Signature: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

 **Overall Grade**

 **Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

**CTS - S2W2**

 **Period: \_\_\_\_\_ Unit: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

 **Grade for this Assignment: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_**

Directions: *This sheet (front and back) is to be completed daily at the beginning of class within the first 5 minutes, signed Thursday night by a parent, and turned in every Monday.*

*Use your notes, vocabulary, and reference materials to answer the questions below.*

#1

***#16 Ammonia gas, NH3 , is stored in a 1000 g bottle. How many moles of NH3 does it contain?***

***#17 What is the mass of 4 mol CO2?***

#2

*\_\_\_\_\_*

***#18 A gas with a pressure of 1 atm and a volume of 4 L is compressed to a new volume of 2 L. What is the***

#3

*\_\_\_\_\_* ***final pressure?***

***#19 How many moles are found in a gas occupying 20 L at STP?***

#4

*\_\_\_\_\_*

***#20 Explain the difference between the terms soluble and insoluble.***

#5

#6

*\_*

***#21 Are gases more or less soluble at high temperatures? What about solids?***

*\_\_*

***#22 Find the molarity of a 2 L solution containing 0.3 mol of HCl.***

#7

***#23 Find the molarity of a 0.6 L solution containing 43 g NaCl.***

#8