**Key Vocabulary: Week 2.5: Periodic Table**

***Complete ALL Week 2.0 and 2.5 flashcards and study over the weekend.***

***Week 2.5- Flashcards on Wednesday AND Thursday, Quiz on Thursday, 6/27.***

1. **Ionic Radii:** Periodic trend that describes the diameter of atoms that have gained or lost electrons according to the octet rule. Increases as you go down a group.
2. **Atomic Radii:** Periodic trend which describes the diameter of normal, uncharged atoms. Increases to the left and down a group.
3. **Electronegativity:** Periodic trend which describes an atom's "pull" on its own electrons as well as those of other atoms. Increases to the right and up a group.
4. **Ionization Energy:** The amount of energy required to remove an electron from an atom. Generally increases up a group and to the right.
5. **Valence Electrons:** Electrons in outer shells of cloud, may be stolen or shared between atoms.
6. **Reactivity:** Periodic trend measuring how likely an element is to react chemically with another element. Generally increases closer to either Francium or Fluorine.
7. **Octet Rule:** The idea that most atoms react to have 8 valence electrons.
8. **Periodic Table:** Chart that organizes the elements into groups and periods by atomic number and valence electrons.
9. **Groups:** Columns "up and down" the periodic table.
10. **Periods:** Rows "left to right" across the periodic table.
11. **Metals:** Elements that are conductive, lustrous, malleable, ductile, and dense. (LEFT)
12. **Non-Metals:** Elements that are poor conductors, dim, and brittle. Mostly gases. (RIGHT)
13. **Semi-Metals/Metalloids:** Elements with properties of both metals and non-metals. (STAIRCASE)