**Key Vocabulary: Week 5.0: Reactions and Review**

***Complete ALL Week 5.0 and 5.5 flashcards and study over the weekend.***

 ***Week 5.0- Flashcards on Monday AND Tuesday, Quiz on Tuesday, 7/16.***

1. **Law of Conservation of Mass:** “What goes in must come out.” The mass of all reactants entering a chemical reaction must match the mass of all products leaving it.
2. **Synthesis/Combination:** A + B 🡪 C Smaller substances combine to make a larger one.
3. **Decomposition:** C 🡪 A + B Larger substances break down to make smaller ones.
4. **Double Replacement:** AB + CD 🡪 AD + BC - A “double swap” where two reactants are rearranged by a chemical reaction.
5. **Single Replacement:** AB + C 🡪 AC +B - A “single swap” where a substance replaces another as part of a compound.
6. **Neutralization:** Acid/base reaction “X(OH) + HY 🡪 H2O + XY” where a substance carrying hydrogen (acid) reacts with a hydroxide base to make water and a salt, XY.
7. **Combustion:** A hydrocarbon reacts with oxygen to produce heat, carbon dioxide, and water. Burning!
8. **Oxidation Numbers:** The charge of an element within a compound.
9. **Precipitation:** Reaction where two dissolved substances react to produce a solid.
10. **Precipitate:** the solid produced when two aqueous substances react.
11. **Coefficient:** large numbers in front of substances in chemical reactions, give the number of moles of that substance.
12. **Balancing Equations:** addition of coefficients to chemical equations to satisfy the law of conservation of mass.