**Key Vocabulary: Week 1.0: Stoichiometry**

***Complete ALL Week 1 and 1.5 flashcards and study over the weekend.***

 ***Week 1.0- Flashcards on Monday AND Tuesday, Quiz on Tuesday, 6/18.***

1. **Atoms:** Smallest particles of matter with the abilities of that matter.
2. **Subscripts:** small number written below chemical symbols in molecules.
3. **Coefficient:** large numbers in front of substances in chemical reactions, give the number of moles of that substance.
4. **Law of Conservation of Mass:** “What goes in must come out.” The mass of all reactants entering a chemical reaction must match the mass of all products leaving it.
5. **Reactants:** Substances reacting or entering a reaction, left side.
6. **Products:** Substances made by a reaction, right side.
7. **Mole:** Unit of chemical measurement. Equal to 602,200,000,000,000,000,000,000 molecules or atoms of a substance.
8. **Avogadro’s Number:** 6.022 x 1023 aka, the number of particles needed to make a mole.
9. **Molar Mass:** mass of a mole of a substance.
10. **Molecular Formula:** the complete formula for a molecule showing the number of atoms of each element it contains.
11. **Empirical Formula:** A molecular formula simplified to show the smallest whole number ratios of atoms. C6H12O6 → CH2O
12. **Gram Atomic Mass:** molar mass of an element or specific type of atom.
13. **Gram Formula Mass:** molar mass of a substance’s empirical formula.